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Excerpt from On the Oxidation of D-Glucose in Alkaline Solution by Air as Well as by Hydrogen Peroxide: A Dissertation Submitted to the Faculty of the Ogden Graduate School of Science in Candidacy for the Degree of Doctor of Philosophy The oxidation of d-glucose in approximately half normal potassium hydroxide solution (9 molecules) by means of hydro gen peroxide has been studied in this laboratory by Spoehr, who found that four oxidation products - namely, small amounts of carbon dioxide (3 51 per glycollic acid and a new hexonic acid, cghuo a-hydroxymethyl-d-arabonic acid, besides very large amounts of formic acid (65 3 per cent.) are formed. About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books. Find more at www.forgottenbooks.com This book is a reproduction of an important historical work. 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Rothberg and Auchinvole (j. Anal. And Appl. Chem. 6, 243 (1892) washed the precipitate with 1% nitric acid solution, then with potassium nitrate to remove the free acid, dissolved it in a standard sodium hydroxide solution and titrated back the excess with nitric acid using phenolphthalein as the indicator. Manby (j. Anal. And Appl. Chem. 6, 82 (1892) and Handy, (j. Anal. And Appl. Chem. 6, 204, (1892) used this same method, standardizing their alkali against the pure precipitate or a steel of known phos phorus content. Pemberton 15, 382 16, 278 (1894) was the first to make an extended investigation of this method. He agrees with Hundeshagen that 23 molecules of alkali are required for one molecule of ammonium phosphomolybdate. Kilgore 16, 765 17, 950 (1895) modified Pem berton's method in that he precipitated the phosphomolybdate at 60c. Instead of at boiling temperature as Pemberton did. Kilgore washed the precipitate with 1% nitric acid solution, then with a 3% potassium nitrate solution and finally with water. He dissolved the precipitate in standard potassium hydroxide solution and titrated back the excess with nitric acid using phenolphthalein as the indi cator. Various other modifications of the pemberton-kilgore method have been suggested. About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books. Find more at www.forgottenbooks.com This book is a reproduction of an important historical work. Forgotten Books uses state-of-the-art technology to digitally reconstruct the work, preserving the original format whilst repairing imperfections present in the aged copy. In rare cases, an imperfection in the original, such as a blemish or missing page, may be replicated in our edition. 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In rare cases, an imperfection in the original, such as a blemish or missing page, may be replicated in our edition. We do, however, repair the vast majority of imperfections successfully; any imperfections that remain are intentionally left to preserve the state of such historical works. Excerpt from Standardization of Potassium Permanganate Solution by Sodium Oxalate The water and sulphuric acid employed were frequently tested for reducing matter and in each case were shown to be free from appreciable amounts Of such impurities. The manganous sulphate was similarly tested for reducing and oxidizing influence and shown to be satisfactory. TO purify the air used for the tests Of Tables VI and VII, it was passed through cotton wool to remove dust and grease, bubbled through a solution Of potassium hydroxide to take out any acid vapors present, and then through, solutions of chromic acid and potassium permanganate. The purified air had no

detectable reducing efiect upon dilute solutions of permanganate under the conditions of titration. The carbon dioxide used for the tests of Table VI was taken from a cylinder of the commercial liquid and passed through water, chromic acid, and permanganate solution before use. An analysis showed the presence of about 3 per cent Of methane; but this gas seemed to have only a very slight reducing action on the permanganate in dilute solution, so that the results of the series in which it was used can be regarded as but slightly less reliable than the series in Table VII. For this latter group Of tests, pure carbon dioxide was made from acid and soda. This source gave a gas which had no detectable reducing action under the conditions Of its use. The potassium permanganate used for most Of the work was a sample Of good quality which had been made up in normal solution for over Six months before filtration and dilution to tenth normal strength for use. The diluted solution was filtered frequently through asbestos to insure freedom from precipitated manganese dioxide. For the series Of tests reported in Table IV, b, a second permanganate was used. In this case the strong solution was boiled for a few minutes, cooled, filtered, and diluted to tenth normal strength. For the series of Table VII still a different permanganate was employed, this stock being prepared in the same manner as the main solution. About 40 grams Of solution were used for each titration. About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books. Find more at www.forgottenbooks.com This book is a reproduction of an important historical work. Forgotten Books uses state-of-the-art technology to digitally reconstruct the work, preserving the original format whilst repairing imperfections present in the aged copy. In rare cases, an imperfection in the original, such as a blemish or missing page, may be replicated in our edition. 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